

Sweet 16 Chemistry Of Gases Tournament

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Sweet 16 Chemistry Of Gases

Sweet 16 Chemistry of Gases Tournament. Do your students eagerly compete to fill out their "March Madness" tournament brackets? Have some fun and inspire your students with March Madness chemistry! This activity combines the popularity of "bracketology" with a review of the preparation and properties of common gases.

Sweet 16 Chemistry of Gases Tournament

An ideal gas adheres exactly to the kinetic theory of gases. 16.3: Pressure: The Result of Constant Molecular Collisions Pressure is a force exerted over an area. Pressure has several common units that can be converted. 16.4: Boyle's Law: Pressure and Volume Boyle's law relates a gas's pressure and volume at constant temperature and amount.

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Read PDF Sweet 16 Chemistry Gases Tournament Answers been provided to help identify the contestants in the Sweet 16 Chemistry of Gases tournament. • First round: Predict the name and formula of the gaseous product obtained in a chemical reaction involving the reactants listed in the first round. The gas "escapes" to the second round!

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Sweet 16 Chemistry Of Gases Tournament Answer Key

First round ----- Predict the gaseous product. 1.) Na₂CO₃ vs. HCl 2.) Zn vs. HCl 3.) H₂O₂ vs. Yeast (catalyst) 4.) Na₂SO₃ vs. HCl 5.) Cu vs. HNO₃ (concentrated) 6.) CaC₂ vs. H₂O 7.) NaOCl vs. HCl 8.) NH₄NO₂ vs. Heat
2nd round- The winner dissolves in H₂O to give an acidic solution. 3rd round (semis) - The gas that is more dense (g/L) wins. Finals - The winner is a greenhouse gas produced in ...

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Sweet 16 Chemistry Compound Tournament Publication No. 10569 With spring just around the corner, your students' thoughts will soon be turning to sunshine, prom, and the NCAA basketball tournament. This clever activity combines the ever-popular March Madness basketball pool with a review of chemical formulas,

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