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2 POWER LAW BASICS

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EQUATION The main

assumption behind the

Arrhenius expression is

that  $k = A \cdot \exp(-E_a/RT)$ . This is

an approximation, but

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it works quite well. The

rate coefficient is the

term that is a function

of temperature but

may also depend

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4.8. The rate of the chemical reaction doubles for an increase of 10 K in absolute temperature from 298 K. Calculate E

a. Ans. 4.9. The activation energy for the reaction,  $2 \text{HI}(\text{g}) \rightarrow \text{H}_2 + \text{I}_2(\text{g})$  is  $209.5 \text{ kJ mol}^{-1}$  at 581 K. Calculate the fraction of molecules of reactants having energy equal to or greater than activation energy?

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Solution The general solution method is best explained in terms of the complex reaction network notation

where the rate for the  $j$ th species is given :  $N$   
 $k_{j1}y_1 k_{j2} y_2 \dots k_{lj} y_j$   
 $\dots dt = \frac{dy_j}{dt} (1)$  A

solution is assumed of the form:  $y_j = Y_j e^{(2)}$

where  $Y_j$  and are

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constants to be  
determined.

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for the detection of a

chemical substance,

that combines a

biological component

with a physicochemical

detector The sensitive

biological element, eg

tissue,

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The reaction definition



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needs information on the kinetics, which was taken from the work by Pereira et al. The heterogeneously catalyzed reaction can be modeled with a Langmuir-Hinshelwood, or...

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