

Chapter 8 Chemistry Covalent Bonding

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Chapter 8 Chemistry Covalent Bonding

1. In covalent bonds, electron sharing usually occurs so that atoms attain the electron configuration of noble gases. (B) 2. Atoms form double or triple covalent bonds if they can attain a noble gas structure by sharing two pairs or three pairs of electrons.

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a covalent bond between atoms in which the electrons are shared unequally a slightly negative charge After attracting electrons more strongly, the more-electronegative atom gains . . .

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A covalent bond in which the bonding electrons are most likely to be found in sausage-shaped regions above and below the bond axis of the bonded atoms Polar molecule A molecule in which one side of the molecule is slightly negative and the opposite side is slightly positive.

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Chapter 8 Covalent Bonding and Molecular Structure 8-3 There are two types of repulsive forces between the two atoms. First, the nuclei repel because they are both positively charged. Second, the electrons repel because they are both negatively charged.

Chapter 8: Covalent Bonding and Molecular Structure

In covalent bonds, atoms share electrons, whereas in ionic bonds, electrons are transferred between atoms. 11. Contrast sigma bonds and pi bonds. A sigma bond is a single covalent bond formed from the direct overlap of orbitals.

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EACH of two atoms provides one electron in a single bond. A dative covalent bond is a special case of covalent bond, in which BOTH electrons in the single bond (bond pair electrons) are provided by the SAME atom. i.e. the donor atom provides both bonding electrons to the acceptor atom in the bond. A dative covalent bond is a covalent bond formed between two atoms where both electrons of the shared pair are contributed by the same atom.

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In covalent bonds, electron sharing usually occurs so that atoms attain the electron configurations of noble gases. • For example, a single hydrogen atom has one electron. But a pair of hydrogen atoms shares electrons to form a covalent bond in a diatomic hydrogen molecule.

Chapter 8

Pure vs. Polar Covalent Bonds. If the atoms that form a covalent bond are identical, as in H₂, Cl₂, and other diatomic molecules, then the electrons in the bond must be shared equally. We refer to this as a pure covalent bond. Electrons shared in pure covalent bonds have an equal probability of being near each nucleus.

Covalent Bonding - Chemistry

Section 8.2 The Nature of Covalent Bonding • OBJECTIVES: –Distinguish between a covalent bond and a coordinate covalent bond, and describe how the strength of a covalent bond is related to its bond dissociation energy.

Chapter 8 "Covalent Bonding"

Covalent Bonding - Chapter 8 1. Covalent Bonding Or How I Learned to Love Sharing (But Remember, File Sharing Is Illegal) 2. As you should remember, ionic compounds are solids at room temperatures that have one ion strip the electron(s) from the other elements' electron cloud. Some compounds do not give up their electrons so easily.

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Chemistry (12th Edition) Chapter 8 - Covalent Bonding ...

A bond formed when two atoms share two electrons. A tightly bonded group of atoms that behaves as a unit and has a negative or positive charge. A neutral group of atoms joined together by covalent bonds. An atom where one side of the molecule is slightly negative and the other is more positive.