

## Chapter 14 Properties Of Gases Answers

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### Chapter 14 Properties Of Gases

states that, at constant volume and temperature, the total pressure exerted by a mixture of gases is equal to the sum of the partial pressure of the component gases diffusion is the tendency of molecules to move toward areas of lower concentration until the concentration is uniform throughout

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Volume. Temperature. Pressure. Amount in Moles. Ideal Gases. considered to be a "point mass". Has elastic collisions ( do not attract or repel each other) Are constant, random straight-line motions. Have an average Kinetic Energy directly related to temperature.

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Chapter 14 The Behavior of Gases147 SECTION 14.1 PROPERTIES OF GASES(pages 413–417) This section uses kinetic theory to explain the properties of gases. This section also explains how gas pressure is affected by the amount of gas, its volume, and its temperature. Compressibility (pages 413–414) 1. Look at Figure 14.1 on page 413.

### SECTION 14.1 PROPERTIES OF GASES(pages 413–417)

14.1 Properties of Gases In organized soccer, there are rules about equipment. For international competitions, the ball's mass must be not more than 450 grams and not less than 410 grams. The pressure of the air inside the ball must be no lower than 0.6 atmospheres and no higher than 1.1 atmospheres at sea level. A ball that is

### 14.1 Properties of Gases 14

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Chapter 14 - The Behavior of Gases - 14.1 Properties of Gases - 14.1 Lesson Check - Page 454: 6 Answer If the temperature is constant, quadrupling the volume would cause the pressure of an enclosed gas to be reduced to one quarter of its original value.

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### **Chapter 14 Properties Of Gases Answers**

CHAPTER 14, The Behavior of Gases (continued) GUIDED PRACTICE PROBLEMS GUIDED PRACTICE PROBLEM 13 (page 424) 13. A gas at 155 kPa and 25°C has an initial volume of 1.00 L. The pressure of the gas increases to 605 kPa as the temperature is raised to 125°C. What is the new volume? SECTION 14.1 PROPERTIES OF GASES(pages 413-417) - MAFIADOC.COM

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Chapter 14- Gases. liquid. Solid. plasma. gas. state in which the particles are held close together, but free.... state in which particles are held close together and can't mov.... state in which particles are far apart, free to flow, and CHAR.... state in which particles are far apart, free to flow and UNCHA....

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### **11.9: The Ideal Gas Law- Pressure, Volume, Temperature ...**

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### **Chapter 14 - The Behavior of Gases - 14.1 Properties of ...**

ST AT 14. small compared to the total volume of the gas. Air will rush into a sealed container when the container is opened. Gas flows from a region of lower pressure to a region of higher pressure. Adding air to an object will cause the object to inflate.

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